# ELECTROVALENT BOND

## GLOSSARY:

- **Anion:** The negatively charged particles.
- **Bond pair:** The pair of electrons shared between two atoms.
- **Cation:** The positively charged particles.
- **Chemical bond:** The attractive force that holds the atoms or ions together in a chemical species.
- **Chemical bonding:** The process, through which chemical bond is formed.
- **Compound:** The combination of two or more elements.
- **Coordinate bond or Dative bond:** The special case of covalent bond formed between two atoms in which both the electrons are contributed by one of the atoms whereas the other atom simply takes part in sharing.
- **Covalent bond:** The bond formed by equal contribution and mutual sharing of electrons between two atoms.
- **Electrovalent or Ionic bond:** The strong electrostatic forces of attraction between oppositely charged ions that hold them together.

- **Inert gas:** A non-reactive gas used during chemical synthesis, chemical analysis or preservation of reactive materials.

- **Ion:** An atom or group of atoms bonded to each other having positive or negative charge and can exist freely in solution.

- **Metallic atoms:** Atoms that donate electrons and gain positive charge.

- **Molecule:** An electrically neutral cluster of mutually bonded atoms.

- **Non-metallic atoms:** Atoms that accept electrons and gain negative charge.

- **Octet rule:** The tendency of an atom of the element to have 8 electrons in its valence shell.

- **Oxidation:** The chemical process, in which an atom or an ion loses one or more electrons.

- **Reduction:** The chemical process, in which an atom or an ion gains one or more electrons.