CALORIMETRY

GLOSSARY:

- **Boiling point:** The fixed temperature at which a liquid gets converted into vapour.
- **Calorimetry:** The measurement of the amount of heat evolved or absorbed in a chemical reaction, change of state or formation of a solution.
- **Condensation:** The process in which vapours, on cooling at a constant temperature, change into liquid.
- **Coolant:** A fluid, flowing through a device to prevent its overheating.
- **Fomentation:** Heating the swollen parts of body at a moderate temperature.
- **Freezing:** The process, where a liquid substance on cooling at a constant temperature, changes into a solid.
- **Freezing point:** The fixed temperature at which a liquid gets converted into solid.
- **Frost:** The solid deposition of water vapor from saturated air.

- **Joule:** The standard unit of energy.
- **Latent heat:** The amount of heat needed to change the state of a given substance without any change in its temperature.
- **Melting:** The process in which a solid, on heating at a constant temperature, changes into liquid.
- **Melting point:** The fixed temperature at which a solid gets converted into liquid.
- **Sublimation:** The change of state of a solid directly into its gaseous state, on heating, without going through the liquid state and vice versa on cooling.
- **Temperature:** The degree of hotness or coldness of a body or environment.
- **Vapourisation:** The process in which a liquid, on heating at a constant temperature, changes into vapour.